

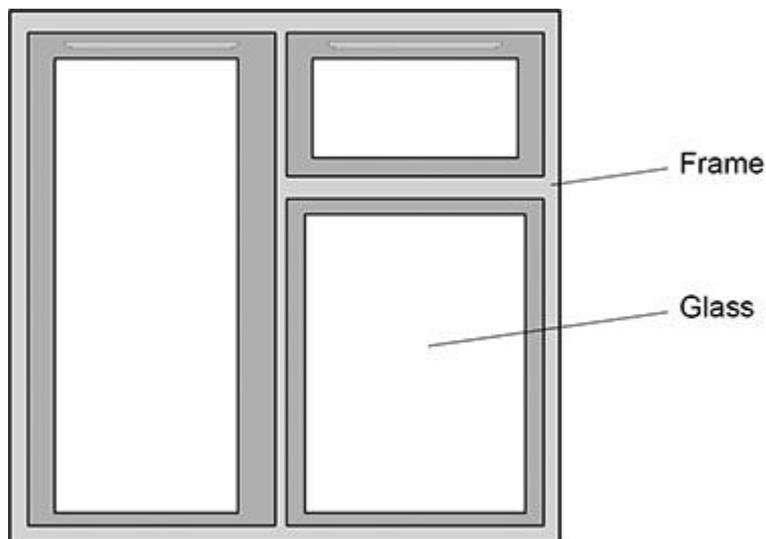
All questions are for separate science students only

Q1.

This question is about substances used to make windows and window frames.

Figure 1 shows a window.

Figure 1



(a) Glass is made by heating sand with two other materials.

Which two other materials are used to make glass?

Tick (✓) two boxes.

Clay

Graphite

Limestone

Sodium carbonate

Sodium hydroxide

(2)

Window frames need to be:

- easy to install
- resistant to damage.

The polymers poly(chloroethene) and HDPE are used to make window frames.

Table 1 shows information about poly(chloroethene) and HDPE.

Table 1

Property	Poly(chloroethene)	HDPE
Density in g/cm ³	1.4	0.92
Relative strength	72	25

- (b) Suggest one advantage of using poly(chloroethene) compared with HDPE to make window frames.
Give one reason for your answer.

Use Table 1.

Advantage _____

Reason _____

(2)

- (c) Suggest one advantage of using HDPE compared with poly(chloroethene) to make window frames.

Give one reason for your answer.

Use Table 1.

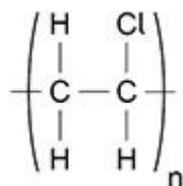
Advantage _____

Reason _____

(2)

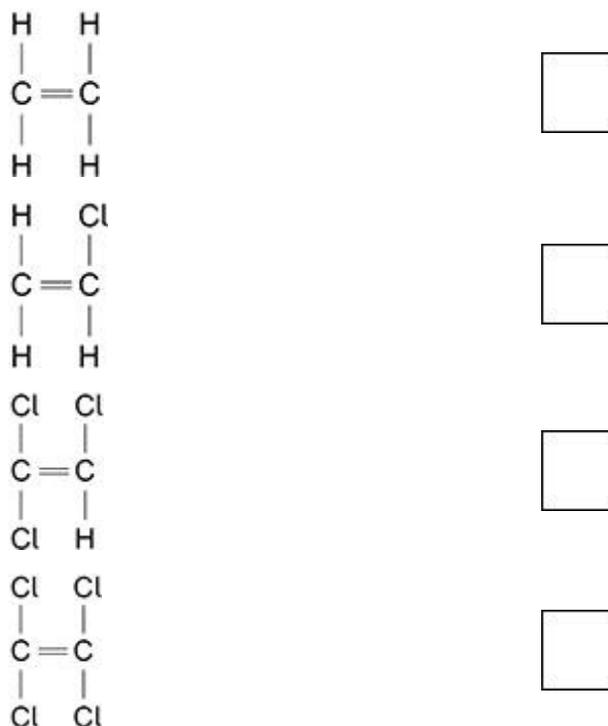
- (d) Figure 2 shows the displayed structural formula of poly(chloroethene).

Figure 2



Which monomer is used to make poly(chloroethene)?

Tick (✓) one box.



(1)

- (e) Chlorine gas is used to produce poly(chloroethene).

Describe a test to identify chlorine gas.

Give the result of the test.

Test _____

Result _____

(2)

- (f) Wood can be used instead of polymers to make window frames.

- Polymers are unreactive.
- Polymers are produced from crude oil.
- Wood breaks down in wet conditions.
- Wood is produced from trees.

Suggest one advantage of using polymers and one advantage of using wood to make window frames.

Advantage of polymers _____

Advantage of wood _____

(2)

Window frames can also be made from an alloy of aluminium.

(g) 6.00 kg of the alloy is used to make a window frame.

Table 2 shows the mass of each element in 6.00 kg of the alloy.

Table 2

Element	Mass in kg
Aluminium	5.94
Magnesium	0.04
Silicon	0.02

Calculate the percentage of aluminium in 6.00 kg of the alloy.

Percentage of aluminium = _____%

(2)

(h) Why is an alloy used instead of pure aluminium to make window frames?

(1)

(Total 14 marks)

Q2.

This question is about organic compounds.

(a) Butane is an alkane with small molecules.

Complete the sentence.

Choose the answer from the box.

fertiliser	formulation	fuel
------------	-------------	------

Butane can be used as a _____.

(1)

(b) Poly(propene) is a polymer.

What is the name of the monomer used to produce poly(propene)?

Tick (✓) one box.

Propane

Propanoic acid

Propanol

Propene

(1)

Ethene and steam react to produce ethanol.

The equation for the reversible reaction is:



- (c) The reaction produces a maximum theoretical mass of 400 kg of ethanol from 243 kg of ethene and 157 kg of steam.
A company produces 380 kg of ethanol from 243 kg of ethene and 157 kg of steam.
The percentage yield of ethanol is less than 100%

Calculate the percentage yield of ethanol.

Use the equation:

$$\text{percentage yield of ethanol} = \frac{\text{mass of ethanol actually made}}{\text{maximum theoretical mass of ethanol}} \times 100$$

Percentage yield = _____ %

(2)

- (d) What are two possible reasons why the percentage yield of ethanol is less than 100%?

Tick (✓) two boxes.

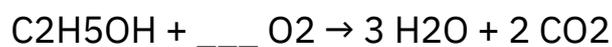
Ethanol is the only product of the reaction.

- Ethanol is very unreactive.
- Some ethanol changes back into ethene and steam.
- Some ethanol escapes from the apparatus.
- Some ethanol reacts with steam.

(2)

- (e) Ethanol burns in oxygen.

Balance the equation for the reaction.



(1)

- (f) Two processes for producing ethanol are:

- fermentation
- hydration (reacting ethene with steam).

The table below shows information about the processes.

Feature	Process	
	Fermentation	Hydration
Raw material	sugar	crude oil
Energy usage	low	high
Rate of reaction	slow	fast
Purity of ethanol	15%	98%

Give two advantages and two disadvantages of using fermentation to produce ethanol.

Advantage of fermentation 1 _____

Advantage of fermentation 2 _____

Disadvantage of fermentation 1 _____

 Disadvantage of fermentation 2 -----

(4)

(Total 11 marks)

Q3.

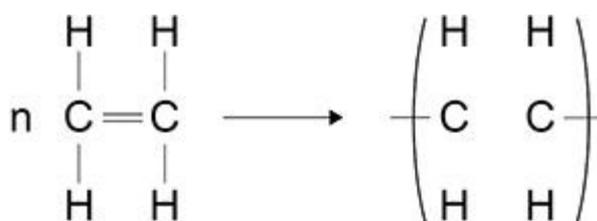
This question is about poly(ethene) and polyesters.

- (a) Poly(ethene) is produced from ethene.

Figure 1 shows part of the displayed structural formula equation for the reaction.

Complete Figure 1.

Figure 1



(2)

- (b) Poly(ethene) is a thermosoftening polymer.

Suggest why poly(ethene) is easier to recycle than thermosetting polymers.

(2)

- (c) Ethene produces different forms of poly(ethene).

How can different forms of poly(ethene) be produced from ethene?

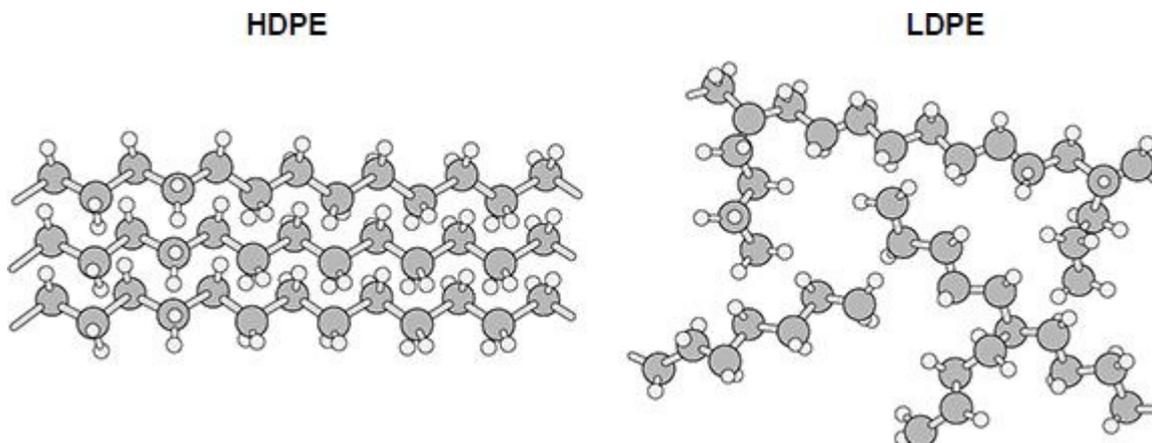
(1)

- (d) Two different forms of poly(ethene) are:

- high density poly(ethene) (HDPE)
- low density poly(ethene) (LDPE).

Figure 2 represents part of the structures of HDPE and LDPE.

Figure 2



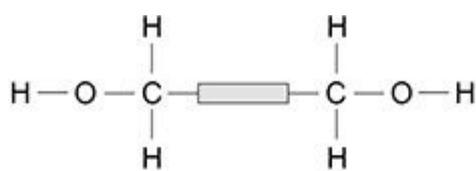
Explain why HDPE has a higher density than LDPE.

(2)

Figure 3 shows three monomers, A, B and C.

Monomer A can react with monomer B and with monomer C to produce polyesters.

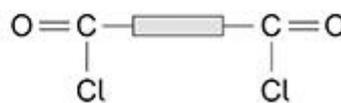
Figure 3



Monomer A



Monomer B



Monomer C

(e) Draw a circle on Figure 3 around an alcohol functional group.

(1)

(f) Complete the table below to show the formula of the small molecule

produced when:

- monomer A reacts with monomer B
- monomer A reacts with monomer C.

Reacting monomers	Formula of small molecule produced
A and B	
A and C	

(1)
(Total 9 marks)

Q4.

This question is about materials used to make plates.

Plates are made from ceramics, paper or poly(propene).

- (a) Paper plates are biodegradable and recyclable.

Which stage of a life cycle assessment (LCA) would contain this information?

Tick (✓) one box.

Disposal at the end of useful life

Extracting and processing raw materials

Manufacturing and packaging

Use and operation during lifetime

(1)

- (b) Which two processes are used to make ceramic plates?

Tick (✓) two boxes.

Forming a composite

Galvanising with zinc

- Heating in a furnace
- Melting sand and boron trioxide
- Shaping wet clay

(2)

Poly(propene) is produced from an alkene.

(c) Complete the sentences.

The name for very large molecules such as poly(propene) is

_____.

The name of the alkene used to produce poly(propene) is

_____.

(2)

(d) The alkene needed to make poly(propene) is produced from crude oil.

Which two processes are used to produce this alkene from crude oil?

Tick (✓) two boxes.

- Chromatography
- Cracking
- Fermentation
- Fractional distillation
- Quarrying

(2)

(e) What type of bond joins the atoms in a molecule of poly(propene)?

Tick (✓) one box.

- Covalent

Ionic	
Metallic	

(1)

The table below shows information about two polymers used to make plates.

Polymer	Effect of heating the polymer
A	does not melt
B	melts at 50 °C

(f) What type of polymer is polymer A?

Use the table above.

(1)

(g) Why does polymer A behave differently to polymer B when heated?

You should refer to crosslinks in your answer.

(1)

(Total 10 marks)

Q5.

This question is about carboxylic acids.

Carboxylic acids belong to a homologous series.

The table below shows information about the first three carboxylic acids in this homologous series.

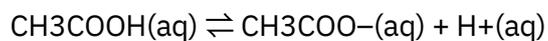
Name	Formula	pH of a 0.01 mol/dm ³ solution
Methanoic acid		2.91
Ethanoic acid	CH ₃ COOH	3.39
	CH ₃ CH ₂ COOH	3.44

(a) Complete the table above.

(2)

- (b) Ethanoic acid ionises in water.

The equation for the reaction is:



Explain how the equation shows that ethanoic acid is a weak acid.

(2)

- (c) A student adds a solution of ethanoic acid to zinc carbonate in an open flask on a balance.

Explain what happens to the mass of the flask and its contents during the reaction.

(3)

- (d) The student compares the rates of the reaction of zinc carbonate with:

- 0.01 mol/dm³ methanoic acid
- 0.01 mol/dm³ ethanoic acid.

The rate of the reaction with methanoic acid is greater than the rate of the reaction with ethanoic acid.

Explain why.

You should refer to ions in your answer.

Use the table above.

(3)

Ethanoic acid reacts with ethanol to produce an ester.

- (e) Give the name of the ester produced when ethanoic acid reacts with ethanol.

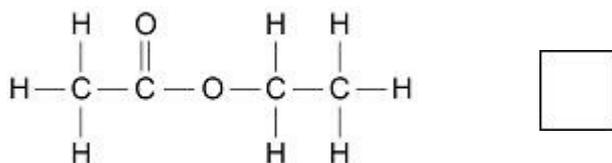
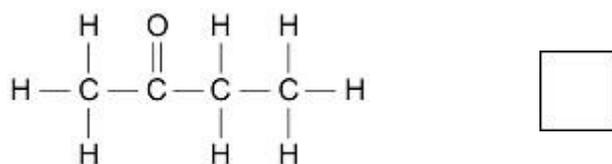
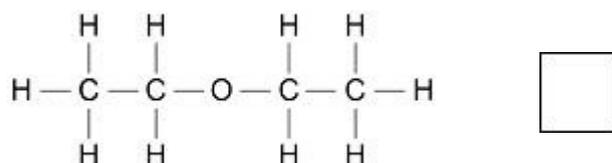
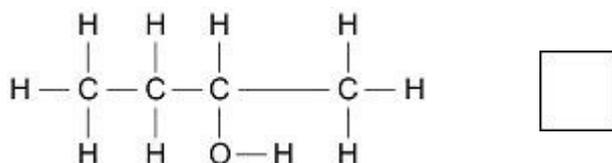
(1)

- (f) Hexanedioic acid and ethanediol join together to produce a polyester.

Ethanoic acid and ethanol join together in the same way to produce an ester.

Which is the displayed structural formula of the ester produced when ethanoic acid reacts with ethanol?

Tick (✓) one box.



(1)

(Total 12 marks)

Q6.

This question is about algae.

A student:

- placed algae in water containing dissolved carbon dioxide
- shone bright light on the algae.

Gas bubbles were collected as the algae photosynthesised.

(a) Describe a test that would identify the gas collected.

Give the result of the test.

Test _____

Result _____

(2)

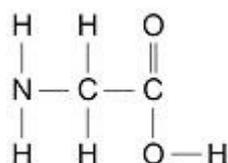
(b) Glucose is produced when algae photosynthesise.

Name two naturally occurring polymers produced from glucose.

_____ and _____

(2)

The diagram below shows the displayed structural formula of an amino acid called glycine.



(c) How many functional groups are there in the molecule in the diagram above?

Tick (✓) one box.

1 2 3 4

(1)

(d) Glycine reacts by condensation polymerisation to produce a polypeptide and one other substance.

Name the other substance produced.

(1)

- (e) Scientists think that algae may have used gases in Earth's early atmosphere.

Algae need an element to produce the molecule in the diagram above which is not present in water or carbon dioxide.

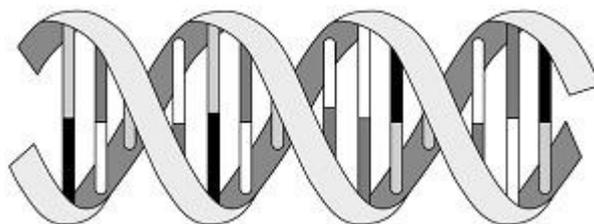
Which two gases from Earth's early atmosphere could have provided this element?

----- and -----

(2)

- (f) The development and function of algae are controlled by a naturally occurring polymer.

The image below represents the shape and structure of this polymer.



Describe the shape and structure of this polymer.

(3)

(Total 11 marks)

Q7.

Figure 1 shows a surfer on a surfboard.

Figure 1



Surfboards are made from polymers.

Surfboards have a poly(styrene) core and an outer skin.

(a) Figure 2 shows the displayed structural formula of poly(styrene).

Figure 2

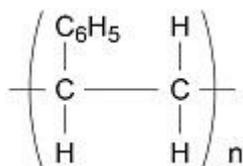
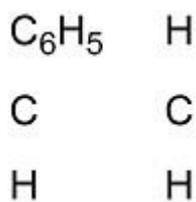


Figure 3 shows an incomplete displayed structural formula of the monomer styrene.

Complete Figure 3.

Figure 3



(2)

The outer skin of surfboards contains a polyester.

Two monomers, A and B, are needed to make the polyester.

Figure 4 shows how these two monomers are represented.

Figure 4



Monomer A



Monomer B

- (b) Name the functional group in monomer B.

----- (1)

- (c) Monomers A and B join together to produce a polyester and a small molecule.

Name the small molecule.

----- (1)

- (d) Why does this type of polyester melt when it is heated?

----- (2)

The outer skin of surfboards is a composite material.

The composite material contains glass fibres surrounded by a polyester.

- (e) Draw one line from each material to the description of that material.

Material	Description of the material
	Hydrocarbon
Glass fibres	Matrix
	Monomer
Polyester	Polypeptide
	Reinforcement

(2)

- (f) The outer skin makes the surfboard more expensive.

Suggest two reasons why an outer skin is added to the poly(styrene) core.

1 _____

2 _____

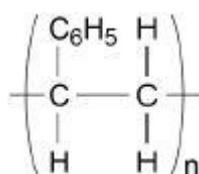
(2)

(Total 10 marks)

Q8.

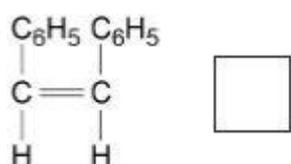
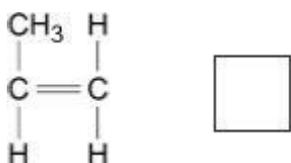
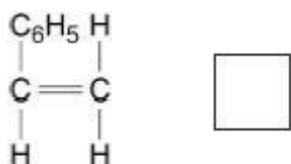
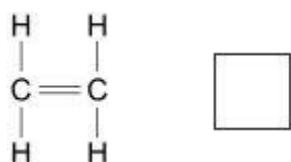
Disposable cups are made from coated paper or poly(styrene).

The diagram below represents the structure of poly(styrene).



- (a) Which small molecule is used to produce poly(styrene)?

Tick one box.



(1)

(b) Which process is used to make poly(styrene) from small molecules?

Tick one box.

Cracking

Distillation

Fermentation

Polymerisation

(1)

(c) Complete the sentences.

Choose answers from the box.

ceramics	composites	four	many
monomers	polymers	two	

Poly(styrene) is produced from small molecules called

This question is about polymers.

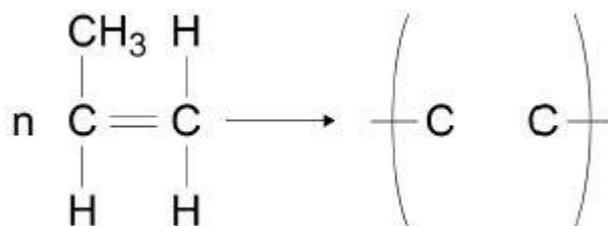
- (a) Polyesters are produced when monomers join together and lose a small molecule.

Name the small molecule lost.

(1)

- (b) Poly(propene) is produced from propene.

Complete the structure of poly(propene) in the equation.



(3)

- (c) Carpets are made from:

- poly(propene)
- wool
- a mixture of poly(propene) and wool.

Poly(propene) wears out more slowly than wool. A mixture of poly(propene) and wool to make carpets is more sustainable than using just poly(propene) or just wool.

Suggest why.

(2)

Polymer fibres are used to make firefighter uniforms.

The table below shows some properties of two polymer fibres.

Polymer fibres

Property	Poly(propene)	Polyester
Density in g/cm ³	0.90	1.38
Melting point in °C	165	260
Flame resistance	Poor	Good
Water absorption	Low	High

- (d) Evaluate the suitability of poly(propene) and polyester for firefighter uniforms.

(4)
(Total 10 marks)

Q10.

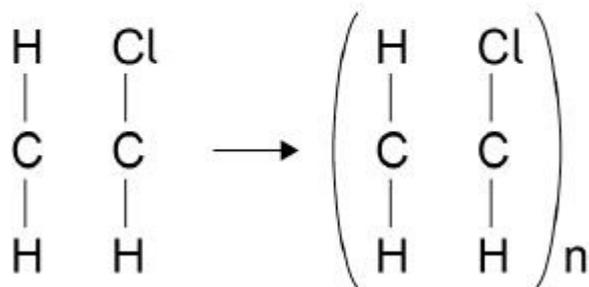
This question is about polymers.

- (a) Name the monomer used to form poly(chloroethene).

(1)

- (b) Figure 1 shows the equation for the formation of poly(chloroethene).
Complete Figure 1.

Figure 1



(3)

- (c) Poly(chloroethene) is the only product.

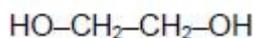
What type of polymer is poly(chloroethene)?

(1)

Ethenediol reacts with butanedioic acid to produce a polyester and a small molecule.

- (d) Figure 2 shows the structural formula of ethenediol.

Figure 2

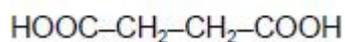


Name the functional group present in ethenediol.

(1)

- (e) Figure 3 shows the structural formula of butanedioic acid.

Figure 3



Which formula represents the carboxylic acid functional group?

Tick (✓) one box.

-CH₂-

-CH₂-CH₂-

-CH₂-COOH

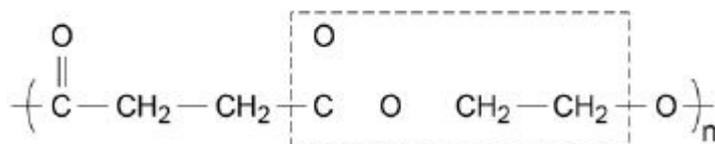


(1)

(f) Figure 4 shows part of the structure of the polyester.

Complete the box in Figure 4.

Figure 4



(2)

(g) Name the small molecule produced when ethanediol reacts with butanedioic acid.

(1)

Starch, proteins and DNA are naturally occurring polymers.

(h) Name the monomers from which starch and proteins are produced.

Starch _____

Proteins _____

(2)

(i) Describe the structure of DNA.

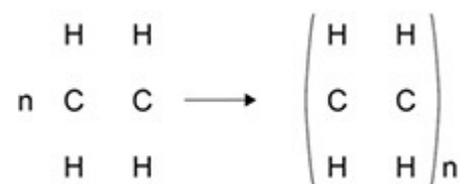
(2)

(Total 14 marks)

Q11.

Ethene is used to produce poly(ethene).

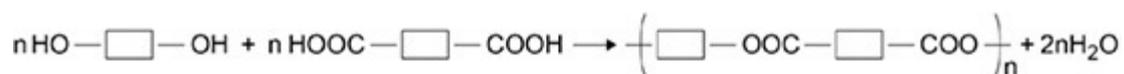
(a) Draw the bonds to complete the displayed formulae of ethene and poly(ethene) in the equation.



(2)

- (b) Polyesters are made by a different method of polymerisation.

The equation for the reaction to produce a polyester can be represented as:



Compare the polymerisation reaction used to produce poly(ethene) with the polymerisation reaction used to produce a polyester.

(4)

(Total 6 marks)