

All questions are for both separate science and combined science students

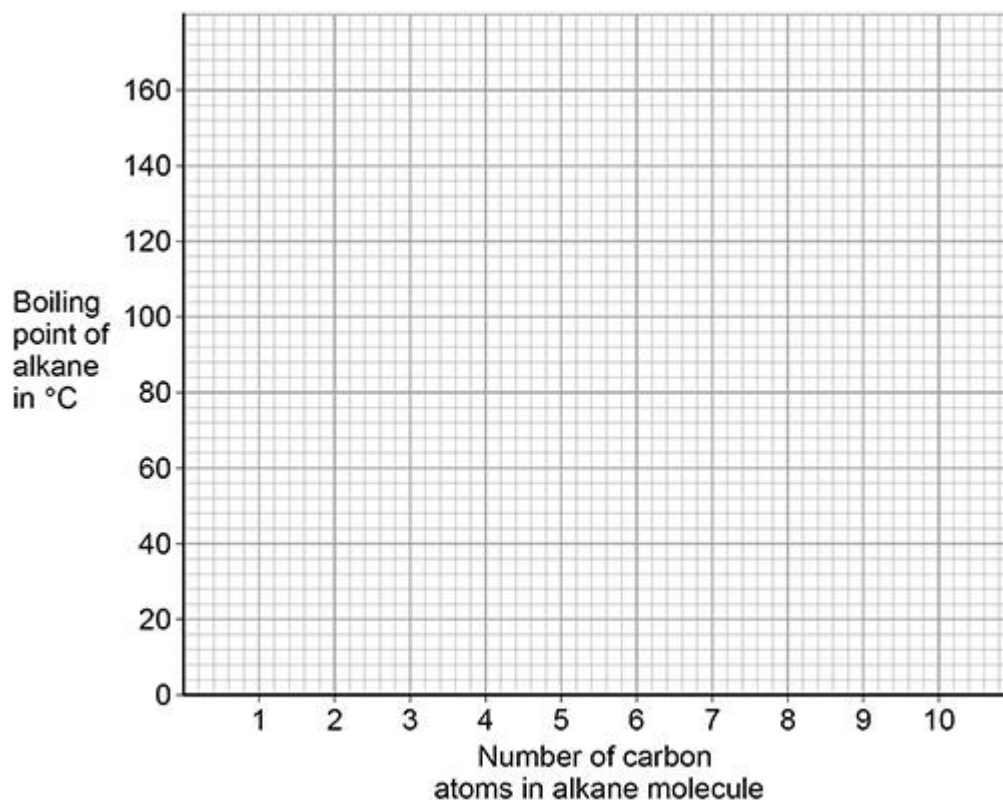
Q1.

This question is about alkanes.

The table below shows information about some alkanes.

Number of carbon atoms in alkane molecule	Boiling point of alkane in °C
4	0
5	36
6	69
7	X
8	126
9	151

(a) Plot the data from the table above on the graph below.



(2)

(b) Predict the boiling point X of the alkane with seven carbon atoms in a molecule.

Use the table and the graph.

X = _____ °C

(1)

- (c) The graph above is not suitable to show the boiling point of the alkane with three carbon atoms in a molecule.

Suggest one reason why.

(1)

- (d) What is the state at 20 °C of the alkane with four carbon atoms in a molecule?

Use the table above.

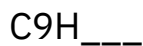
(1)

The table in part (a) is repeated below.

Number of carbon atoms in alkane molecule	Boiling point of alkane in °C
4	0
5	36
6	69
7	X
8	126
9	151

The alkane with nine carbon atoms in a molecule is called nonane.

- (e) Complete the formula of nonane.



(1)

- (f) Nonane will condense lower in a fractionating column during fractional distillation than the other alkanes in the table above.

Explain why.

You should refer to the temperature gradient in the fractionating column.

(2)
(Total 8 marks)

Q2.

This question is about hydrocarbons.

Hexane and hexene are hydrocarbons containing six carbon atoms in each molecule.

Hexane is an alkane and hexene is an alkene.

(a) Draw one line from each hydrocarbon to the formula of that hydrocarbon.

Hydrocarbon

Formula

	C6H8
Hexane	C6H10
	C6H12
Hexene	C6H14
	C6H16

(2)

(b) Bromine water is added to hexane and to hexene.

What would be observed when bromine water is added to hexane and to hexene?

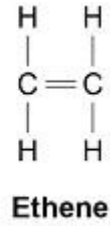
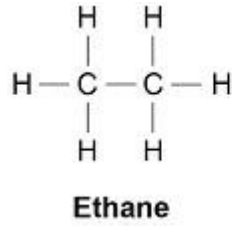
Hexane _____

Hexene _____

(2)

(c) Ethane is an alkane and ethene is an alkene.

The diagram below shows the displayed structural formulae of ethane and of ethene.



Compare ethane with ethene.

You should refer to:

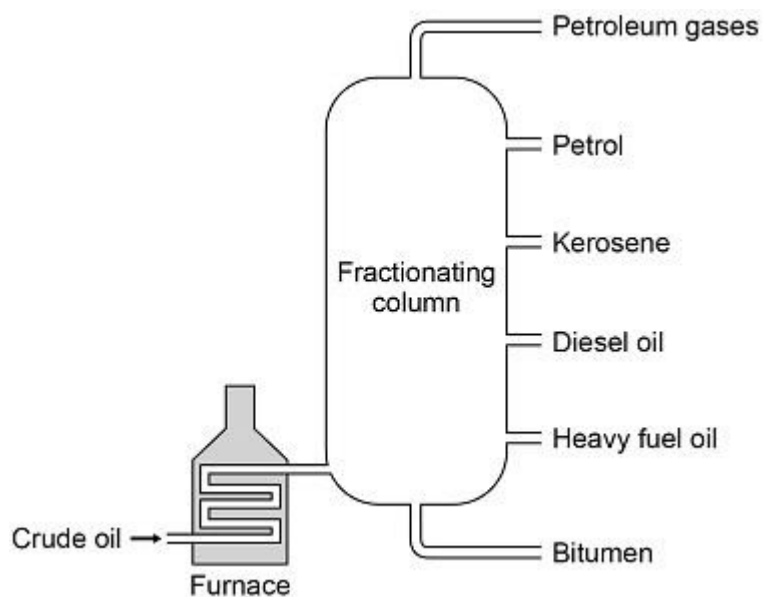
- their structure and bonding
- their reactions.

(6)
(Total 10 marks)

Q3.

This question is about crude oil and hydrocarbons.

The figure below shows a fractionating column used to separate crude oil into fractions.



The following table gives information about some of the fractions.

Fraction	Boiling point range in °C
Petroleum gases	Below 30
Petrol	40-110
Kerosene	180-260
Diesel oil	260-320
Heavy fuel oil	320-400
Bitumen	400-450

(a) Suggest a suitable temperature for the furnace in the figure.

_____ °C

(1)

(b) Explain why diesel oil collects above heavy fuel oil but below kerosene in the fractionating column.

Use the table above.

(2)

(c) Suggest two reasons why bitumen is not used as a fuel.

1 _____

2 _____

(2)

(d) Petrol contains mainly alkanes.

Which of the following compounds is an alkane?

Tick (✓) one box.

C₂H₄

C₄H₈

C₆H₁₄

C₈H₁₆

(1)

Large hydrocarbon molecules in the diesel oil fraction are cracked to produce smaller hydrocarbon molecules.

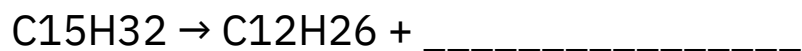
(e) Describe the conditions needed to crack hydrocarbon molecules from the diesel oil fraction.

(2)

(f) Explain why large hydrocarbon molecules in the diesel oil fraction are cracked to produce smaller hydrocarbon molecules.

(2)

- (g) Complete the equation for the cracking of C
- ₁₅
- H
- ₃₂



(1)

(Total 11 marks)

Q4.

This question is about combustion of fuels.

- (a) Some central heating boilers use wood as a fuel.

Suggest two reasons why wood is more sustainable than natural gas as a fuel for central heating boilers.

1 -----

2 -----

(2)

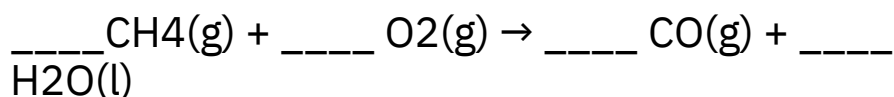
Natural gas is mainly methane.

When methane burns it can produce both carbon monoxide and carbon dioxide.

- (b) Explain the process by which carbon monoxide can be produced when methane is burned.

(2)

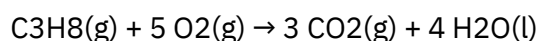
- (c) Balance the equation for the combustion of methane to produce carbon monoxide.



(1)

- (d) Propane burns to form carbon dioxide and water.

The equation for the reaction is:



3.60 dm³ carbon dioxide is produced when a sample of propane is burned in 7.25 dm³ oxygen.

Calculate the volume of unreacted oxygen.

Give your answer in cm³

Volume of unreacted oxygen = _____ cm³

(4)

(Total 9 marks)

Q5.

This question is about fuels.

Octane (C₈H₁₈) is a hydrocarbon in petrol.

- (a) Cracking breaks down large hydrocarbon molecules into smaller hydrocarbon molecules.

Which hydrocarbon molecule can be cracked to produce octane, C₈H₁₈?

Tick one box.

C₄H₈

C₄H₁₀

C₈H₁₆

C₁₂H₆

(1)

- (b) What type of carbon compound is octane, C₈H₁₈?

Tick one box.

Alcohol

Alkane

Carboxylic acid

Ester

(1)

(c) Oxygen is needed to burn fuels.

Name the source of the oxygen needed to burn fuels.

(1)

(d) Particulates and sulfur dioxide are pollutants produced when some fuels burn.

Draw one line from each pollutant to the polluting effect.

Pollutant	Polluting effect
	Acid rain
Particulates	Global dimming
	Global warming
Sulfur dioxide	Landfill
	Sewage sludge

(2)

(e) Which two gases are produced when fuels burn in car engines?

Tick two boxes.

Ammonia

Carbon dioxide	<input type="checkbox"/>
Carbon monoxide	<input type="checkbox"/>
Nitrogen	<input type="checkbox"/>
Oxygen	<input type="checkbox"/>

(2)

(f) Vehicles produce most of the atmospheric pollution in cities.

How could the atmospheric pollution in cities be reduced?

Tick two boxes.

Build more roads in cities	<input type="checkbox"/>
Build new car factories	<input type="checkbox"/>
Develop fuel efficient engines	<input type="checkbox"/>
Make car tax cheaper	<input type="checkbox"/>
Use electric cars	<input type="checkbox"/>

(2)

(Total 9 marks)

Q6.

This question is about hydrocarbons.

The table gives information about four hydrocarbons.

The hydrocarbons are four successive members of a homologous series.

Hydrocarbon	Formula	Boiling point in °C
A	C ₄ H ₁₀	0
B		36
C	C ₆ H ₁₄	69

D	C ₇ H ₁₆	98
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(a) What is the formula of hydrocarbon B?

Tick (✓) one box.

C₄H₁₂

C₅H₁₂

C₅H₁₀

C₆H₁₂

(1)

(b) What is the simplest ratio of carbon : hydrogen atoms in a molecule of hydrocarbon A?

Ratio = 2 : _____

(1)

(c) Which hydrocarbon is a gas at room temperature (25 °C)?

Tick (✓) one box.

A B C D

(1)

(d) Which hydrocarbon is most flammable?

Tick (✓) one box.

A B C D

(1)

(e) Which two substances are produced when a hydrocarbon completely combusts in air?

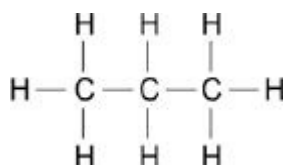
Tick (✓) two boxes.

Carbon

Carbon dioxide	<input type="checkbox"/>
Hydrogen	<input type="checkbox"/>
Sulfur dioxide	<input type="checkbox"/>
Water	<input type="checkbox"/>

(2)

The diagram shows the displayed structure of a hydrocarbon molecule.



(f) What is the name of the hydrocarbon in the diagram above?

Tick (✓) one box.

Butane	<input type="checkbox"/>
Ethane	<input type="checkbox"/>
Methane	<input type="checkbox"/>
Propane	<input type="checkbox"/>

(1)

(g) Calculate the relative formula mass (M_r) of the hydrocarbon in the diagram above.

Relative atomic masses (A_r): H = 1 C = 12

Relative formula mass (M_r) = -----

(2)

(Total 9 marks)

Q7.

This question is about alkenes and crude oil.

- (a) Pentene is an alkene molecule containing five carbon atoms.

Complete the formula for pentene.

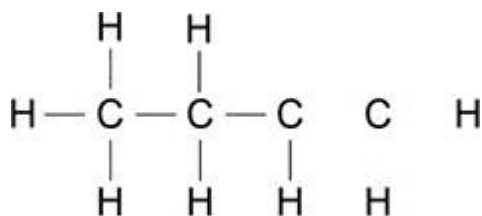


(1)

- (b) Butene is an alkene molecule containing four carbon atoms.

The diagram shows all of the atoms and some of the bonds in the displayed formula for butene.

Complete the displayed formula by adding the remaining bonds.



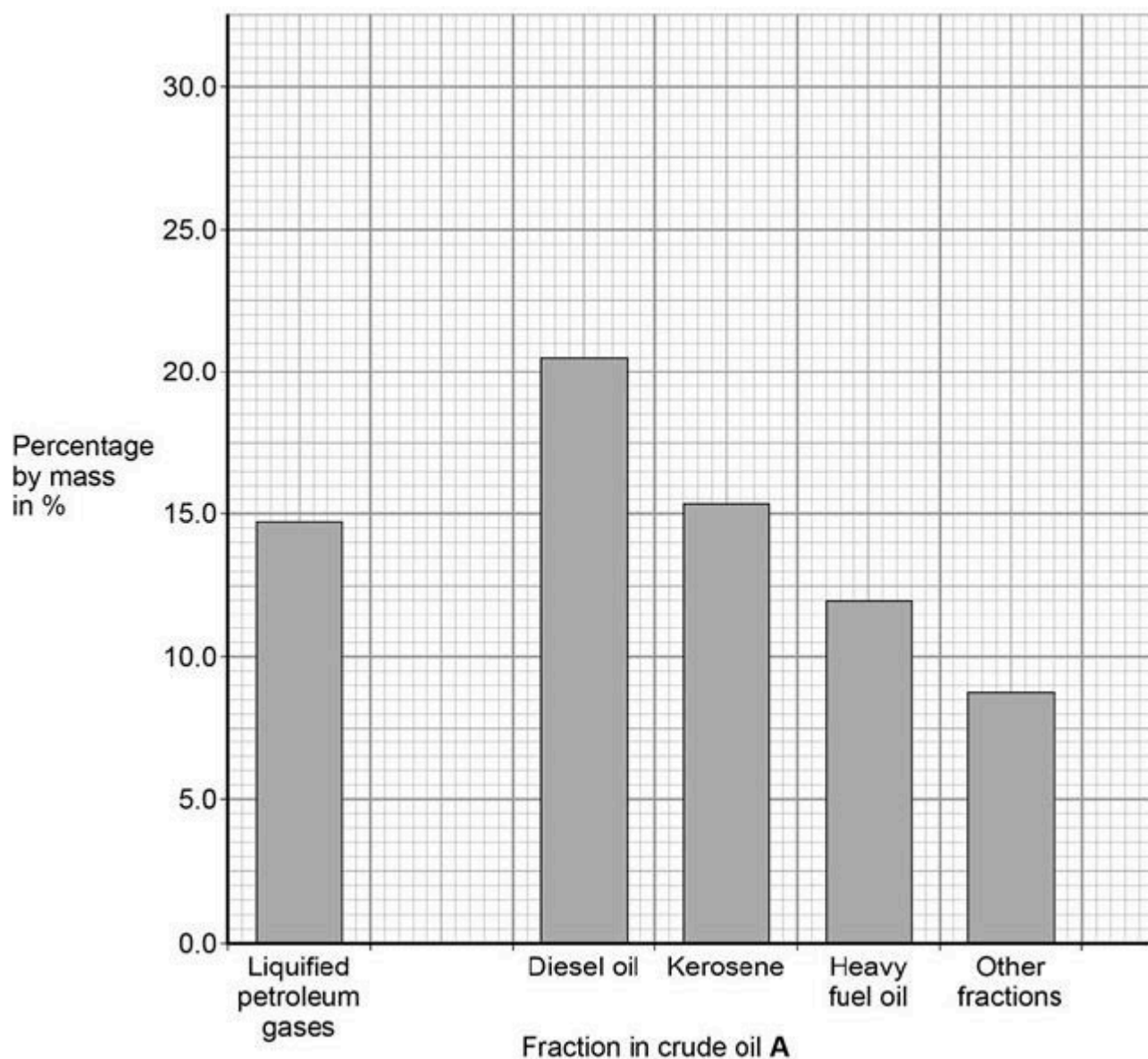
(1)

Pentene and butene are produced from crude oil.

The table shows the percentages of different fractions in two samples of crude oil.

Fraction	Percentages by mass in %	
	Crude oil A	Crude oil B
Liquefied petroleum gases	14.7	7.1
Petrol	28.6	11.1
Diesel oil	20.5	17.2
Kerosene	15.4	38.5
Heavy fuel oil	12.0	16.0
Other fractions	8.8	10.1

The graph shows the percentages of different fractions in crude oil A.



(c) Plot the data for petrol in the table above on the graph.

(1)

(d) What mass of crude oil A is needed to obtain 12 tonnes of heavy fuel oil?.

Use the table above.

10 tonnes

100 tonnes

1000 tonnes

10 000 tonnes



(1)

- (e) What mass of crude oil A is needed to obtain 12 tonnes of heavy fuel oil?.

Calculate the total mass of car fuel that can be produced from 2000 kg of crude oil B.

Use the table above.

Mass of car fuel = _____ kg

(3)

- (f) Crude oil B is a better source of hydrocarbons for cracking than crude oil A. Suggest why.

Use the table above.

(1)

- (g) Alkenes are obtained from crude oil using fractional distillation followed by cracking.

Explain how alkenes are produced using fractional distillation followed by cracking.

(6)

(Total 14 marks)

Q8.

This question is about hydrocarbons.

- (a) The names and formulae of three hydrocarbons in the same homologous series are:

Ethane	C ₂ H ₆
Propane	C ₃ H ₈
Butane	C ₄ H ₁₀

The next member in the series is pentane.

What is the formula of pentane?

(1)

(b) Which homologous series contains ethane, propane and butane?

Tick one box.

Alcohols

Alkanes

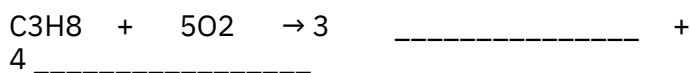
Alkenes

Carboxylic acids

(1)

(c) Propane (C₃H₈) is used as a fuel.

Complete the equation for the complete combustion of propane.



(2)

(d) Octane (C₈H₁₈) is a hydrocarbon found in petrol. Explain why octane is a hydrocarbon.

(2)

(e) The table below gives information about the pollutants produced by cars using diesel or petrol as a fuel.

Fuel	Relative amounts of pollutants		
	Oxides of Nitrogen	Particulate matter	Carbon dioxide
Diesel	31	100	85
Petrol	23	0	100

Compare the pollutants from cars using diesel with those from cars using petrol.

(3)

(f) Pollutants cause environmental impacts.

Draw one line from each pollutant to the environmental impact caused by the pollutant.

Pollutant	Environmental impact caused by the pollutant
	Acid rain
Oxides of nitrogen	Flooding
	Global dimming
Particulate matter	Global warming
	Photosynthesis

(2)

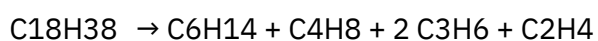
(Total 11 marks)

Q9.

This question is about organic compounds.

Hydrocarbons can be cracked to produce smaller molecules.

The equation shows the reaction for a hydrocarbon, C₁₈H₃₈



(a) Which product of the reaction shown is an alkane?

Tick one box.

C₂H₄C₃H₆C₄H₈C₆H₁₄

(1)

- (b) The table below shows the boiling point, flammability and viscosity of C₁₈H₃₈ compared with the other hydrocarbons shown in the equation.

	Boiling point	Flammability	Viscosity
A	highest	lowest	highest
B	highest	lowest	lowest
C	lowest	highest	highest
D	lowest	highest	lowest

Which letter, A, B, C or D, shows how the properties of C₁₈H₃₈ compare with the properties of C₂H₄, C₃H₆, C₄H₈ and C₆H₁₄?

Tick one box.

A

B

C

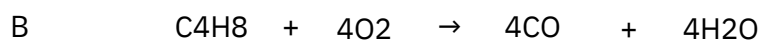
D

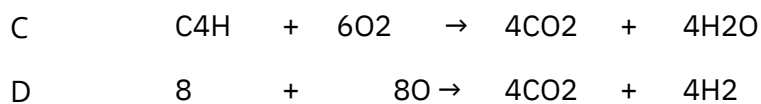
(1)

- (c) The hydrocarbon C₄H₈ was burnt in air.

Incomplete combustion occurred.

Which equation, A, B, C or D, correctly represents the incomplete combustion reaction?





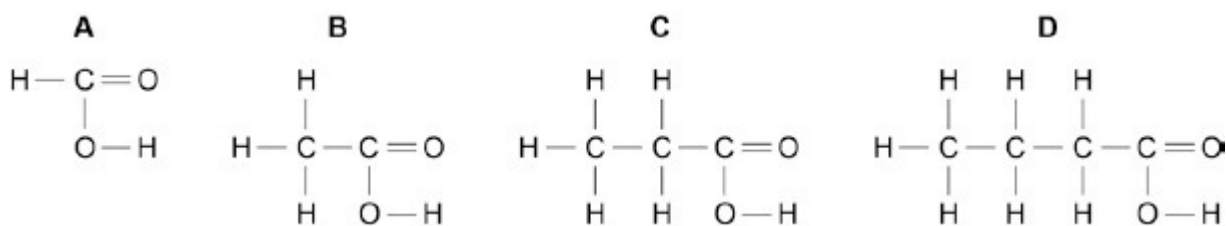
Tick one box.

A	^{C₄H₈} 8	<input type="checkbox"/>
B		<input type="checkbox"/>
C		<input type="checkbox"/>
D		<input type="checkbox"/>

(1)

(d) Propanoic acid is a carboxylic acid.

Which structure, A, B, C or D, shows propanoic acid?



Tick one box.

A	<input type="checkbox"/>
B	<input type="checkbox"/>
C	<input type="checkbox"/>
D	<input type="checkbox"/>

(1)

(e) Propanoic acid is formed by the oxidation of which organic compound?

Tick one box.

Propane	<input type="checkbox"/>
Propene	<input type="checkbox"/>

Propanol

Polyester

(1)

(Total 5 marks)

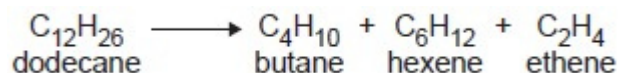
Q10.

This question is about hydrocarbons.

(a) Most of the hydrocarbons in crude oil are alkanes.

(i) Large alkane molecules can be cracked to produce more useful molecules.

The equation shows the cracking of dodecane.



Give two conditions used to crack large alkane molecules. 1.

..... 2.

.....

(2)

(ii) The products hexene and ethene are alkenes.

Complete the sentence.

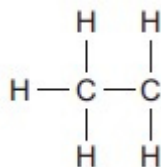
When alkenes react with bromine water the colour changes

from orange to

(1)

(iii) Butane (C₄H₁₀) is an alkane.

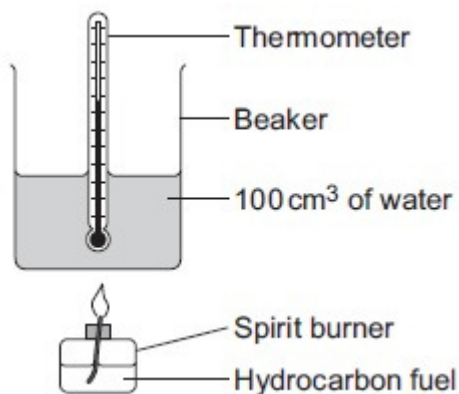
Complete the displayed structure of butane.



(1)

(b) A group of students investigated the energy released by the combustion of four hydrocarbon fuels.

The diagram below shows the apparatus used.



Each hydrocarbon fuel was burned for two minutes.

Table 1 shows the students' results.

Table 1

Name and formula of hydrocarbon fuel	After two minutes			Energy released by 1.0 g of fuel in kJ	Relative amount of smoke in the flame
	Mass of fuel used in g	Temperature increase of water in °C	Energy released by fuel in kJ		
Hexane, C ₆ H ₁₄	0.81	40	16.80	20.74	very little smoke
Octane, C ₈ H ₁₈	1.10	54	22.68	20.62	some smoke
Decane, C ₁₀ H ₂₂	1.20	58	24.36		smoky
Dodecane, C ₁₂ H ₂₆	1.41	67	28.14	19.96	very smoky

- (i) Calculate the energy released by 1.0 g of decane in kJ.

Energy released = _____ kJ

(2)

- (ii) Suggest one improvement to the apparatus, or the use of the apparatus, that would make the temperature increase of the water for each fuel more accurate.

Give a reason why this is an improvement.

----- (2)

(iii) The students noticed that the bottom of the beaker became covered in a black substance when burning these fuels.

Name this black substance.

Suggest why it is produced.

(2)

(iv) A student concluded that hexane is the best of the four fuels.
Give two reasons why the results in Table 2 support this conclusion.

1. -----

2. -----

(2)

(c) In this question you will be assessed on using good English, organising information clearly and using specialist terms where appropriate.

Most car engines use petrol as a fuel.

- Petrol is produced from the fractional distillation of crude oil.
- Crude oil is a mixture of hydrocarbons.
- Sulfur is an impurity in crude oil.

Car engines could be developed to burn hydrogen as a fuel.

- Hydrogen is produced from natural gas.
- Natural gas is mainly methane.

Table 2 shows information about petrol and hydrogen.

Table 2

Petrol	Hydrogen
--------	----------

State of fuel at room temperature	Liquid	Gas
Word equation for combustion of the fuel	petrol + oxygen →	hydrogen + oxygen →
Energy released from combustion of 1 g of the fuel	carbon dioxide + water 47 kJ	water 142 kJ

Describe the advantages and disadvantages of using hydrogen instead of petrol in car engines.

Use the information given and your knowledge and understanding to answer this question.

(6)

(Total 18 marks)