All questions are for separate science students only

Q1.

This question is about salts.

(a) Name the salt produced by the neutralisation of hydrochloric acid with potassium hydroxide.

_____

(1)

(b) Write an ionic equation for the neutralisation of hydrochloric acid with potassium hydroxide.

_____ + _____ → _____

(1)

(c) Soluble salts can be produced by reacting dilute hydrochloric acid with an insoluble solid.

Copper, copper carbonate and copper oxide are insoluble solids.

Which of these insoluble solids can be used to make a copper salt by reacting the solid with dilute hydrochloric acid?

Tick (✓) one box.

Copper and copper carbonate only	☐
Copper and copper oxide only	☐
Copper carbonate and copper oxide only	☐
Copper, copper carbonate and copper oxide	☐

(1)

A student makes crystals of magnesium sulfate.

This is the method used.

1. Add sulfuric acid to a beaker.
2. Warm the sulfuric acid.
3. Add a spatula of magnesium oxide to the beaker.
4. Stir the mixture.
5. Repeat steps 3 and 4 until there is magnesium oxide remaining in the beaker.
6. Filter the mixture.
7. Evaporate the filtrate gently until crystals start to form.

8. Leave the solution to finish crystallising.

(d) Give one reason for:

- step 2
- step 5
- step 6.

Step 2 _____

_____

Step 5 _____

_____

Step 6 _____

_____

(3)

(e) How should the filtrate be evaporated gently in step 7?

_____

_____

(1)

(f) Iron chloride is produced by heating iron in chlorine gas.

The equation for the reaction is:

$$2\ Fe + 3\ Cl_2 \rightarrow 2\ FeCl_3$$

Calculate the volume of chlorine needed to react with 14 g of iron.

You should calculate:

- the number of moles of iron used
- the number of moles of chlorine that react with 14 g of iron
- the volume of chlorine needed.

Relative atomic mass (*Ar*): Fe = 56

The volume of 1 mole of gas = 24 dm3

_____

_____

_____

_____

_____

_____

Volume of chlorine = _____dm3

(3)

(Total 10 marks)

Q2.

This question is about the reaction between hydrogen sulfide (H2S) and oxygen.

The equation for the reaction is:

$$2\ H_2S(g) + 3\ O_2(g) \rightarrow 2\ H_2O(g) + 2\ SO_2(g)$$

(a) What does H2O(g) represent?

_____

(1)

(b) Calculate the volume of oxygen required to react with 50 cm3 of hydrogen sulfide.

_____

_____

Volume = _____cm3

(1)

(c) Figure 1 shows part of the reaction profile for the reaction.

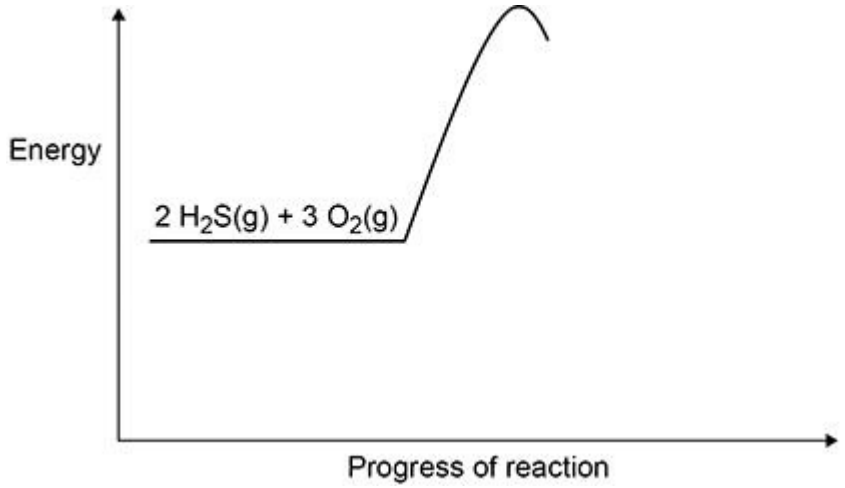
The reaction is exothermic.

Complete Figure 1.

You should:

- complete the profile line
- label the activation energy
- label the overall energy change.

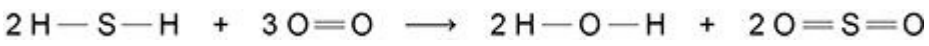
Figure 1



(3)

(d) Figure 2 shows the displayed formula equation for the reaction of hydrogen sulfide with oxygen.

Figure 2



The table below shows some of the bond energies.

Bond	H—S	O=O	H—O	S=O
Energy in kJ/mol	364	498	464	X

In the reaction the energy released forming new bonds is 1034 kJ/mol greater than the energy needed to break existing bonds.
Calculate the bond energy X for the bond.

Use Figure 2 and the table above.

X = _____kJ/mol

(5)

(Total 10 marks)

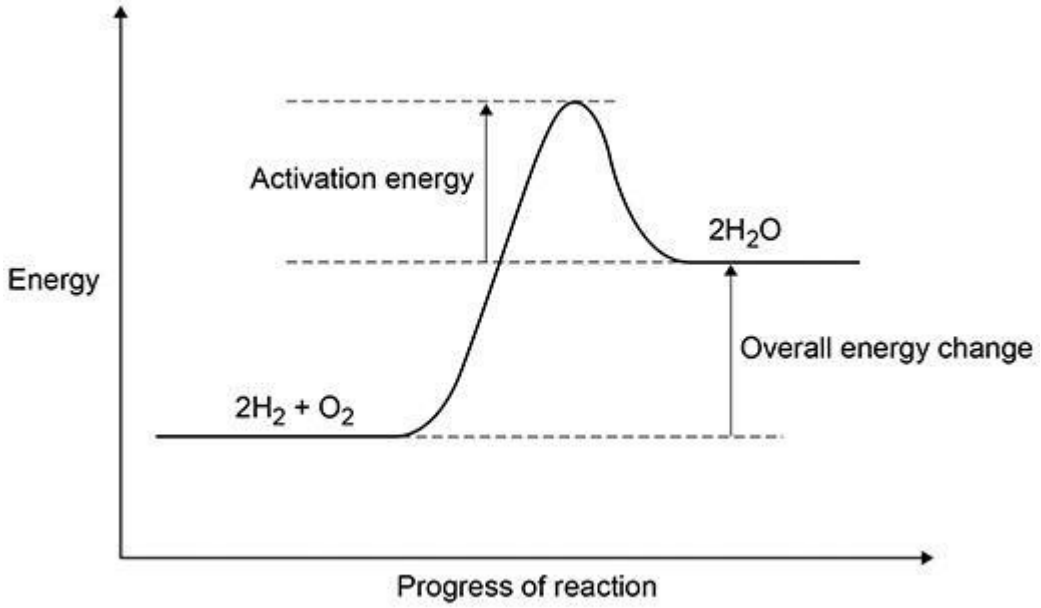
Q3.

The reaction between hydrogen and oxygen releases energy.

(a) A student drew a reaction profile for the reaction between hydrogen and oxygen.

Figure 1 shows the student's reaction profile.

Figure 1



The student made two errors when drawing the reaction profile.

Describe the two errors.

1 _____

_____

2 _____

_____

(2)

(b) The reaction between hydrogen and oxygen in a hydrogen fuel cell is used to produce electricity.

Hydrogen fuel cells and rechargeable cells are used to power some cars.

Give two advantages of using hydrogen fuel cells instead of using rechargeable cells to power cars.

1 _____

_____

2 _____

_____

(2)

(c) Reactions occur at the positive electrode and at the negative electrode in a hydrogen fuel cell.

Write a half equation for one of these reactions.

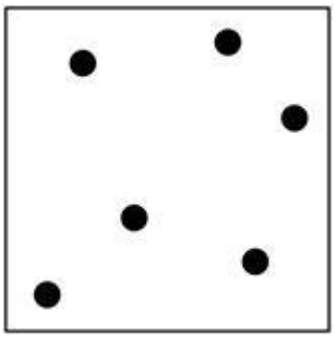
_____

(1)

(d) The three states of matter can be represented by a simple particle model.

Figure 2 shows a simple particle model for hydrogen gas.

Figure 2



Give two limitations of this simple particle model for hydrogen gas.

1 _____

_____

2 _____

_____

(2)

(e) The hydrogen gas needed to power a car for 400 km would occupy a large volume.

Suggest one way that this volume can be reduced.

_____

_____

(1)

(f) The energy needed for a car powered by a hydrogen fuel cell to travel 100 km is 58 megajoules (MJ).

The energy released when 1 mole of hydrogen gas reacts with oxygen is 290 kJ

The volume of 1 mole of a gas at room temperature and pressure is 24 dm3

Calculate the volume of hydrogen gas at room temperature and pressure needed for the car to travel 100 km

_____

_____

_____

_____

_____

_____

_____

_____

Volume of hydrogen gas = _____ dm3

(4)

(Total 12 marks)

Q4.

This question is about electrolysis.

Aluminium is produced by electrolysing a molten mixture of aluminium oxide and cryolite.

(a) Explain why a mixture is used as the electrolyte instead of using only aluminium oxide.

_____

_____

_____

_____

(2)

(b) What happens at the negative electrode during the production of aluminium?

Tick (✓) one box.

Aluminium atoms gain electrons. ☐

Aluminium atoms lose electrons. ☐

Aluminium ions gain electrons. ☐

Aluminium ions lose electrons. ☐

(1)

(c) Oxygen is produced at the positive electrode.

Complete the balanced half-equation for the process at the positive electrode.

→ O2 +

(2)

(d) Explain why the positive electrode must be continually replaced.

_____

_____

_____

_____

_____

_____

(3)

(e) The overall equation for the electrolysis of aluminium oxide is:

$$2\ Al_2O_3 \rightarrow 4\ Al + 3\ O_2$$

Calculate the mass of oxygen produced when 2000 kg of aluminium oxide is completely electrolysed.

Relative atomic masses (*Ar*): O = 16 Al = 27

_____

_____

_____

_____

_____

_____

_____

_____

Mass of oxygen = _____ kg

(4)

Sodium metal and chlorine gas are produced by the electrolysis of molten sodium chloride.

(f) Explain why sodium chloride solution cannot be used as the electrolyte to produce sodium metal.

_____

_____

_____

_____

(2)

(g) Calculate the volume of 150 kg of chlorine gas at room temperature and pressure.

The volume of one mole of any gas at room temperature and pressure is 24.0 dm3

Relative formula mass (*Mr*): Cl2 = 71

_____

_____

_____

_____

Volume = _____ dm3

(2)

(Total 16 marks)

Q5.

This question is about combustion of fuels.

(a) Some central heating boilers use wood as a fuel.

Suggest two reasons why wood is more sustainable than natural gas as a fuel for central heating boilers.

1 _____

_____

2 _____

_____

(2)

Natural gas is mainly methane.

When methane burns it can produce both carbon monoxide and carbon dioxide.

(b) Explain the process by which carbon monoxide can be produced when methane is burned.

_____

_____

_____

_____

(2)

(c) Balance the equation for the combustion of methane to produce carbon monoxide.

____$CH_4(g)$ + ____ $O_2(g)$ → ____ $CO(g)$ + ____ $H_2O(l)$

(1)

(d) Propane burns to form carbon dioxide and water.

The equation for the reaction is:

$$C_3H_8(g) + 5\ O_2(g) \rightarrow 3\ CO_2(g) + 4\ H_2O(l)$$

3.60 dm^3 carbon dioxide is produced when a sample of propane is burned in 7.25 dm^3 oxygen.

Calculate the volume of unreacted oxygen.

Give your answer in cm^3

_____

_____

_____

_____

_____

_____

_____

_____

Volume of unreacted oxygen = _____ cm3

(4)

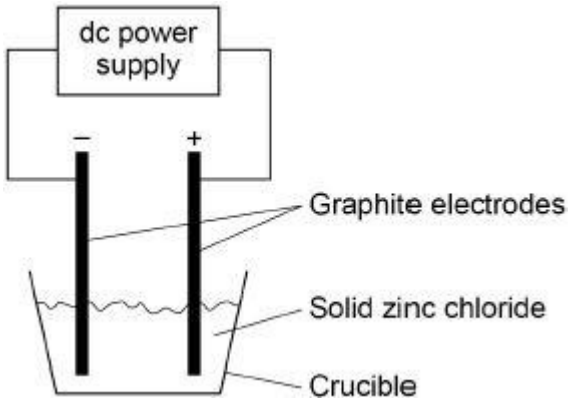
(Total 9 marks)

Q6.

A student investigated the electrolysis of different substances.

Figure 1 shows the apparatus.

Figure 1



(a) Explain why electrolysis would not take place in the apparatus shown in Figure 1.

_____

_____

_____

_____

(2)

(b) Explain why graphite conducts electricity.

Answer in terms of the structure and bonding in graphite.

_____

_____

_____

_____

_____

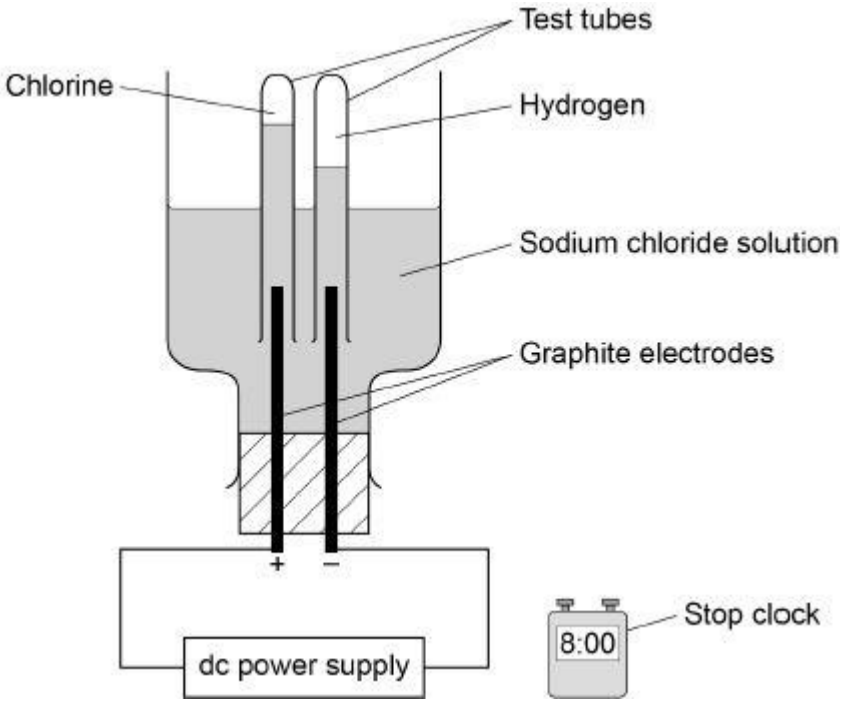
_____

(3)

The student investigated how the volume of gases produced changes with time in the electrolysis of sodium chloride solution.

Figure 2 shows the apparatus.

Figure 2



(c) The student made an error in selecting the apparatus for this investigation.

How should the apparatus be changed? Give one reason for your answer. _____

_____

_____

_____

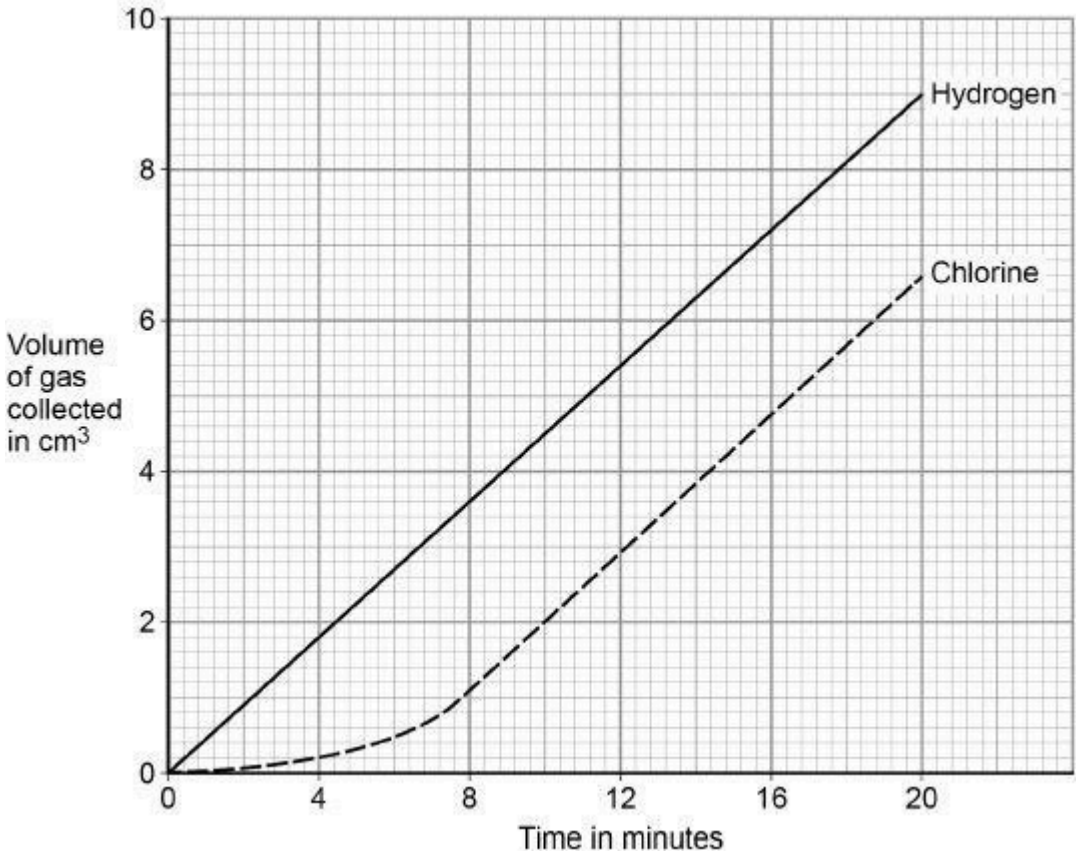
(2)

Another student used the correct apparatus.

This student measured the volumes of gases collected every minute for 20 minutes.

Figure 3 shows the student's results.

Figure 3



(d) Describe the trends shown in the results.

Use values from Figure 3.

_____

_____

_____

_____

_____

_____

(3)

(e) The number of moles of each gas produced at the electrodes is the same.

No gas escapes from the apparatus. Suggest one reason for the difference in volume of each gas collected.

_____

_____

(1)

(f) Calculate the amount in moles of chlorine collected after 20 minutes.

Use Figure 3.

The volume of one mole of any gas at room temperature and pressure is 24.0 dm3

Give your answer in standard form.

_____

_____

_____

_____

Moles of chlorine = _____ mol

(3)

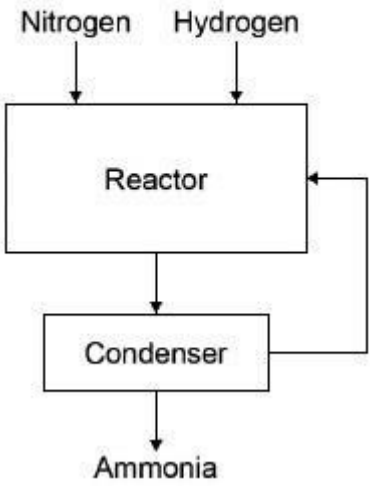
(Total 14 marks)

Q7.

Nitrogen and hydrogen react to produce ammonia in the Haber process.

Figure 1 shows the Haber process.

Figure 1



A gaseous mixture of ammonia, hydrogen and nitrogen leaves the reactor.

Table 1 shows the boiling points of the gases.

Table 1

Gas	Boiling point in °C
Ammonia	−33
Nitrogen	−196

Hydrogen	−253

(a) Suggest how ammonia is separated from the other gases.

_____

_____

_____

_____

_____

(2)

(b) What happens to the unreacted hydrogen and nitrogen?

_____

_____

(1)

The equation for the reaction is:

$$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$$

The forward reaction is exothermic.

(c) Calculate the volume of ammonia produced from the complete reaction of 825 dm^3 of hydrogen.

_____

_____

_____

_____

Volume of ammonia = _____ dm^3

(2)

(d) The Haber process uses a temperature of 450 °C and a pressure of 200 atmospheres.

Why are these conditions used?

Tick two boxes.

A higher pressure is maintained using less energy ☐

A higher temperature would increase the equilibrium yield ☐

A lower pressure would decrease the equilibrium yield ☐

A lower temperature would make the reaction too slow ☐

There are more product molecules than reactant molecules ☐

(2)

Most of the ammonia produced is used to make fertilisers.

Table 2 shows information about compounds used as fertilisers.

Table 2

Compound	Formula	Cost in £ / tonne
A	NH4NO3	220
B	(NH4)2HPO4	350
C	KCl	235

(e) Which element in compound A improves agricultural productivity?

_____

(1)

(f) Which two compounds can be mixed to make a fertiliser containing three elements that improve agricultural productivity?

Give a reason why you have chosen these compounds.

Compounds _____ and _____

Reason

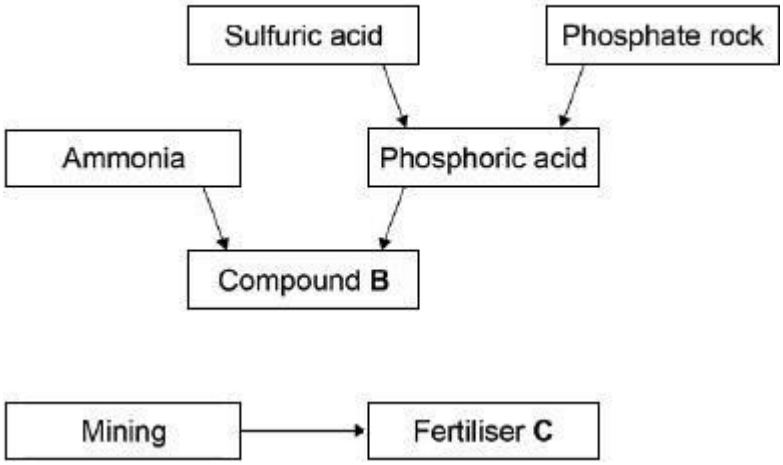
_____

_____

(2)

(g) Figure 2 shows a flow chart for the production of compounds B and C.

Figure 2



Suggest two possible reasons for the difference in cost between compounds B and C.

1. _____

_____

2. _____

_____

(2)

(Total 12 marks)