All questions are for separate science students only

Q1.

This question is about the extraction of metals.

Element R is extracted from its oxide by reduction with hydrogen.

The equation for the reaction is:

$$3\ H_2 + RO_3 \rightarrow R + 3\ H_2O$$

(a) The sum of the relative formula masses (*Mr*) of the reactants ($3\ H_2 + RO_3$) is 150

Calculate the relative atomic mass (*Ar*) of R.

Relative atomic masses (*Ar*): H = 1 O = 16

_____

_____

_____

_____

Relative atomic mass (*Ar*) of R = _____

(2)

(b) Identify element R.

You should use:

- your answer to part (a)
- the periodic table.

Identity of R = _____

(1)

(c) Carbon is used to extract tin (Sn) from tin oxide (SnO_2).

The equation for the reaction is:

$$SnO_2 + C \rightarrow Sn + CO_2$$

Calculate the percentage atom economy for extracting tin in this reaction.

Relative atomic masses (*Ar*): C = 12 O = 16 Sn = 119

_____

_____

_____

_____

_____

_____

Percentage atom economy = _____ %

(3)

(d) Tungsten (W) is a metal.

Tungsten is extracted from tungsten oxide (WO3).

All other solid products from the extraction method must be separated from the tungsten.

The table below shows information about three possible methods to extract tungsten from tungsten oxide.

Method	Reactant	Relative cost of reactant	Products
1	Carbon	Low	Tungsten solid Carbon dioxide gas Tungsten carbide solid
2	Hydrogen	High	Tungsten solid Water vapour
3	Iron	Low	Tungsten solid Iron oxide solid

Evaluate the three possible methods for extracting tungsten from tungsten oxide.

_____

_____

_____

_____

_____

_____

_____

_____

_____

_____

_____

_____

(4)

(Total 10 marks)

Q2.

This question is about organic compounds.

(a) Butane is an alkane with small molecules.

Complete the sentence.

Choose the answer from the box.

fertiliser	formulation	fuel

Butane can be used as a _____.

(1)

(b) Poly(propene) is a polymer.

What is the name of the monomer used to produce poly(propene)?

Tick (✓) one box.

Propane ☐

Propanoic acid ☐

Propanol ☐

Propene ☐

(1)

Ethene and steam react to produce ethanol.

The equation for the reversible reaction is:

ethene + steam ⇌ ethanol

(c) The reaction produces a maximum theoretical mass of 400 kg of ethanol from 243 kg of ethene and 157 kg of steam.

A company produces 380 kg of ethanol from 243 kg of ethene and 157 kg of steam.

The percentage yield of ethanol is less than 100%

Calculate the percentage yield of ethanol.

Use the equation:

$$\text{percentage yield of ethanol} = \frac{\text{mass of ethanol actually made}}{\text{maximum theoretical mass of ethanol}} \times 100$$

_____

_____

Percentage yield = _____ %

(2)

(d) What are two possible reasons why the percentage yield of ethanol is less than 100%?

Tick (✓) two boxes.

Ethanol is the only product of the reaction.	☐
Ethanol is very unreactive.	☐
Some ethanol changes back into ethene and steam.	☐
Some ethanol escapes from the apparatus.	☐
Some ethanol reacts with steam.	☐

(2)

(e) Ethanol burns in oxygen.

Balance the equation for the reaction.

$$C_2H_5OH + \_\_\_ O_2 \rightarrow 3\ H_2O + 2\ CO_2$$

(1)

(f) Two processes for producing ethanol are:

- fermentation
- hydration (reacting ethene with steam).

The table below shows information about the processes.

Feature	Process	
	Fermentation	Hydration

Raw material	sugar	crude oil
Energy usage	low	high
Rate of reaction	slow	fast
Purity of ethanol	15%	98%

Give two advantages and two disadvantages of using fermentation to produce ethanol.

Advantage of fermentation 1 _____

_____

Advantage of fermentation 2 _____

_____

Disadvantage of fermentation 1 _____

_____

Disadvantage of fermentation 2 _____

_____

(4)

(Total 11 marks)

Q3.

This question is about silver iodide.

Silver iodide is produced in the reaction between silver nitrate solution and sodium iodide solution.

The equation for the reaction is:

$$AgNO3(aq) + NaI(aq) \rightarrow AgI(s) + NaNO3(aq)$$

(a) A student investigated the law of conservation of mass.

This is the method used.

1. Pour silver nitrate solution into a beaker labelled A.
2. Pour sodium iodide solution into a beaker labelled B.
3. Measure the masses of both beakers and their contents.
4. Pour the solution from beaker B into beaker A.
5. Measure the masses of both beakers and their contents again.

The table below shows the student's results.

	Mass before mixing in g	Mass after mixing in g
Beaker A and contents	78.26	108.22
Beaker B and contents	78.50	48.54

Explain how the results demonstrate the law of conservation of mass.

You should use data from table above in your answer.

_____

_____

_____

_____

(2)

(b) Suggest how the student could separate the insoluble silver iodide from the mixture at the end of the reaction.

_____

_____

(1)

The student purified the separated silver iodide.

This is the method used.

1. Rinse the silver iodide with distilled water.
2. Warm the silver iodide.

(c) Suggest one impurity that was removed by rinsing with water.

_____

_____

(1)

(d) Suggest why the student warmed the silver iodide.

_____

_____

(1)

(e) Calculate the percentage atom economy for the production of silver iodide in this reaction.

The equation for the reaction is:

$$AgNO_3(aq) + NaI(aq) \rightarrow AgI(s) + NaNO_3(aq)$$

Give your answer to 3 significant figures.

Relative formula masses:

(*M*r): $AgNO_3 = 170$ NaI = 150 AgI = 235 $NaNO_3 = 85$

_____

_____

_____

_____

_____

_____

_____

_____

Percentage atom economy (3 significant figures) = _____ %

(4)

(f) Give one reason why reactions with a high atom economy are used in industry.

_____

_____

(1)

(Total 10 marks)

Q4.

This question is about the extraction of metals.

(a) Tungsten is a metal.

The symbol of tungsten is W

Tungsten is produced from tungsten oxide by reaction with hydrogen.

The equation for the reaction is:

$$WO_3 + 3\ H_2 \rightarrow W + 3\ H_2O$$

Calculate the percentage atom economy when tungsten is produced in this reaction.

Use the equation:

$$\text{percentage atom economy} = \frac{184}{(M_r\ WO_3) + (3 \times M_r\ H_2)} \times 100$$

Relative formula masses (*M*r): WO3 = 232 H2 = 2

_____

_____

_____

_____

Percentage atom economy = _____%

(2)

Aluminium is extracted from aluminium oxide.

(b) 38% of a rock sample is aluminium oxide. Calculate the mass of aluminium oxide in 40 kg of the rock sample.

_____

_____

_____

_____

Mass of aluminium oxide = _____ kg

(2)

(c) The formula of aluminium oxide is Al2O3

Calculate the relative formula mass (*Mr*) of aluminium oxide.

Relative atomic masses (*Ar*): O = 16 Al = 27

_____

_____

_____

_____

Relative formula mass (*Mr*) = _____

(2)

(d) 60.0 kg of aluminium oxide produces a maximum of 31.8 kg of aluminium.

In an extraction process only 28.4 kg of aluminium is produced from 60.0 kg of aluminium oxide.

Calculate the percentage yield.

Give your answer to 3 significant figures.

Use the equation:

$$\text{percentage yield} = \frac{\text{mass of product actually made}}{\text{maximum theoretical mass of product}} \times 100$$

_____

_____

_____

_____

_____

_____

_____

Percentage yield = _____%

(3)

(e) Extracting metals by electrolysis is a very expensive process.

Explain why aluminium is extracted using electrolysis and not by reduction with carbon.

_____

_____

_____

_____

(2)

(Total 11 marks)

Q5.

This question is about reversible reactions and equilibrium.

Hydrogen is used to produce ammonia in the Haber process.

The hydrogen is made in two stages.

Stage 1 is the reaction of methane and steam to produce carbon monoxide and hydrogen.

The equation for the reaction is:

$$CH_4(g) + H_2O(g) \rightleftharpoons CO(g) + 3\,H_2(g)$$

(a) Calculate the atom economy for the formation of hydrogen in stage 1.

Relative atomic masses (*Ar*): H = 1 C = 12 O = 16

_____

_____

_____

_____

Atom economy = _____%

(2)

(b) Explain why a low pressure is used in stage 1.

Give your answer in terms of equilibrium.

_____

_____

_____

_____

(2)

(c) Stage 2 uses the carbon monoxide produced in stage 1.

The carbon monoxide is reacted with more steam to produce carbon dioxide and more hydrogen.

The equation for the reaction in stage 2 is:

$$CO(g) + H2O(g) \rightleftharpoons CO2(g) + H2(g)$$

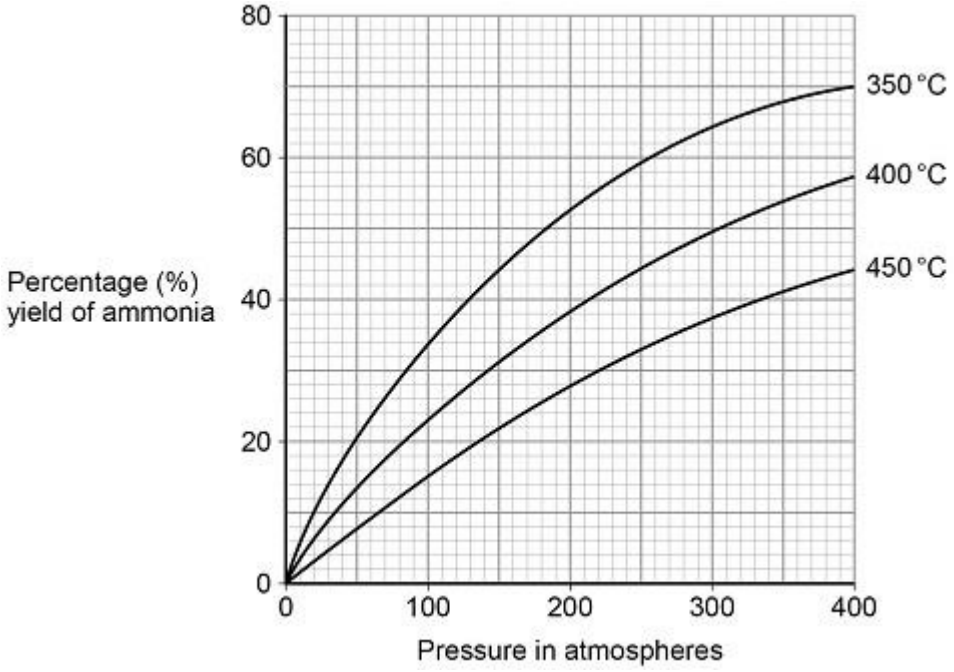
What is the effect of increasing the pressure on the equilibrium yield of hydrogen in stage 2?

_____

_____

(1)

The graph below shows the percentage yield of ammonia produced at different temperatures and pressures in the Haber process.



A temperature of 450 °C and a pressure of 200 atmospheres are used in the Haber process.

(d) A student suggested that a temperature of 350 °C and a pressure of 285 atmospheres could be used instead of those used in the Haber process.
Determine how many times greater the percentage yield of ammonia obtained would be.
Use the graph.

_____

_____

_____

_____

Percentage yield = _____ times greater

(3)

(e) A pressure of 285 atmospheres is not used in the Haber process instead of 200 atmospheres.

Give one reason why.

_____

_____

(1)

(f) How does the graph above show that the forward reaction in the Haber process is exothermic?

_____

_____

(1)

(g) World production of ammonia is now about 30 times greater than it was in 1950.

Suggest why the demand for ammonia has increased.

_____

_____

_____

_____

(2)

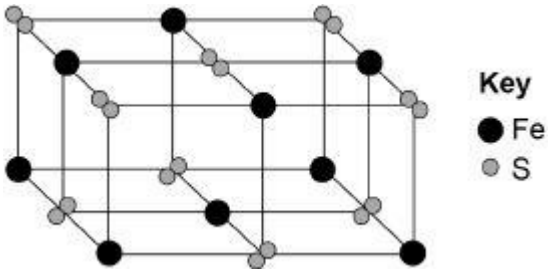
(Total 12 marks)

Q6.

This question is about metals and metal compounds.

(a) Iron pyrites is an ionic compound.

The diagram below shows a structure for iron pyrites.



Determine the formula of iron pyrites.

Use the diagram above.

_____

(1)

(b) An atom of iron is represented as $^{56}_{26}Fe$

Give the number of protons, neutrons and electrons in this atom of iron.

Number of protons ___ Number of neutrons

_____ Number of electrons

_____

(3)

(c) Iron is a transition metal.

Sodium is a Group 1 metal.

Give two differences between the properties of iron and sodium.

1. _____

_____

2. _____

_____

(2)

Nickel is extracted from nickel oxide by reduction with carbon.

(d) Explain why carbon can be used to extract nickel from nickel oxide. _____

_____

_____

_____

(2)

(e) An equation for the reaction is:

$$NiO + C \rightarrow Ni + CO$$

Calculate the percentage atom economy for the reaction to produce nickel.

Relative atomic masses (*Ar*): C = 12 Relative formula mass (*Mr*): NiO = 75 Give your answer to 3 significant figures.

_____

_____

_____

_____

_____

_____

Percentage atom economy = _____ %

(3)

(Total 11 marks)

Q7.

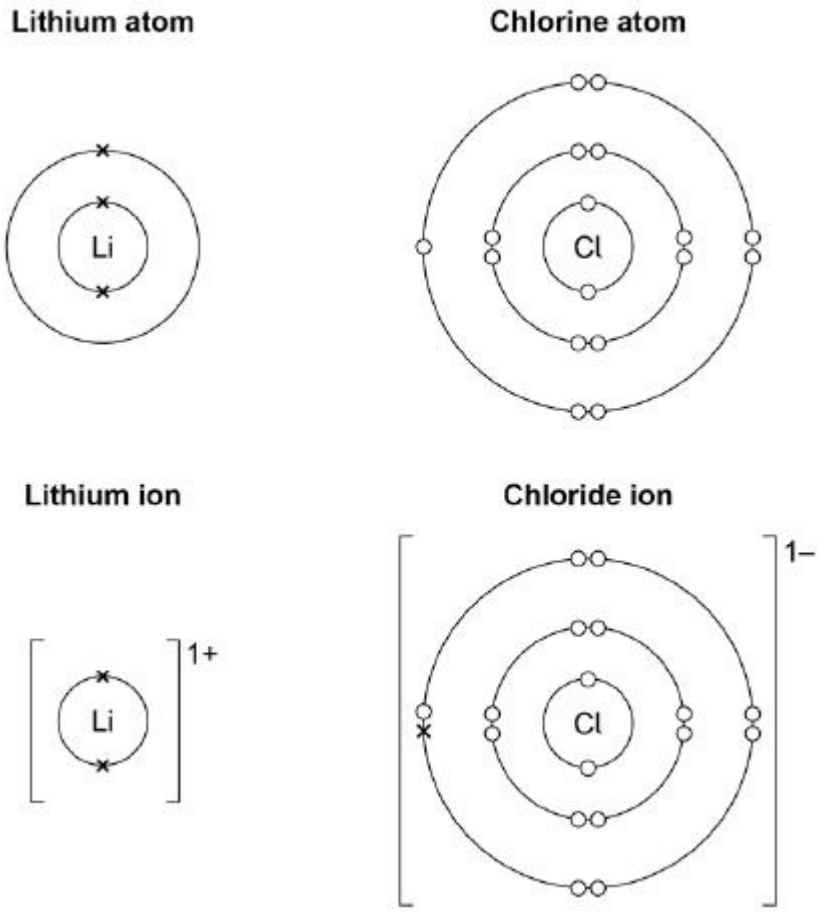
This question is about metal compounds.

(a) Lithium reacts with chlorine to produce lithium chloride.

When lithium atoms and chlorine atoms react to produce lithium chloride, lithium ions and chloride ions are formed.

The diagram shows the electronic structures of the atoms and ions.

The symbols o and x are used to represent electrons.



Describe what happens when a lithium atom reacts with a chlorine atom. Answer in terms of electrons.

_____

_____

_____

_____

_____

_____

_____

_____

(4)

Zinc sulfate can be made by two methods.

The equations for the two methods are:

Method 1: ZnO + H2SO4 → ZnSO4 + H2O

Method 2: ZnCO3 + H2SO4 → ZnSO4 + H2O + CO2

(b) Calculate the percentage atom economy for making zinc sulfate in Method 1.

Use the equation:

percentage atom economy =

$$\frac{\text{relative formula mass of } ZnSO_4}{\text{relative formula mass of ZnO + relative formula mass of } H_2SO_4} \times 100$$

Give your answer to 3 significant figures.

Relative formula masses (*Mr*): ZnO = 81 H2SO4 = 98 ZnSO4 = 161

_____

_____

_____

_____

_____

Percentage atom economy = _____ %

(3)

(c) Method 1 gives a higher percentage atom economy for making zinc sulfate than Method 2.

Give a reason why it is important to use a reaction with a high atom economy.

_____

_____

(1)

(d) A student uses 50 cm3 of a zinc sulfate solution of 80 g/dm3

What mass of zinc sulfate is dissolved in 50 cm3 of this zinc sulfate solution?

_____

_____

_____

_____

Mass = _____ g

(2)

(Total 10 marks)

Q8.

A scientist produces zinc iodide (ZnI_2).

This is the method used.

1. Weigh 0.500 g of iodine.
2. Dissolve the iodine in ethanol.
3. Add an excess of zinc.
4. Stir the mixture until there is no further change.
5. Filter off the excess zinc.
6. Evaporate off the ethanol.

(a) Ethanol is flammable.

Suggest how the scientist could carry out Step 6 safely.

_____

_____

(1)

(b) Explain why the scientist adds excess zinc rather than excess iodine. _____

_____

_____

_____

_____

_____

(3)

(c) Calculate the minimum mass of zinc that needs to be added to 0.500 g of iodine so that the iodine fully reacts.

The equation for the reaction is:

$$Zn + I_2 \rightarrow ZnI_2$$

Relative atomic masses (*Mr*): Zn = 65 I = 127

_____

_____

_____

_____

_____

_____

Minimum mass of zinc = _____ g

(3)

A different scientist makes zinc iodide by the same method.

The scientist obtains 12.5 g of zinc iodide.

The percentage yield in this reaction is 92.0%.

(d) What is the maximum theoretical mass of zinc iodide produced in this reaction?

_____

_____

_____

_____

_____

_____

Maximum theoretical mass = _____ g

(3)

(e) Suggest one reason why the percentage yield in this reaction is not 100%.

_____

_____

(1)

(f) The scientist makes a solution of zinc iodide with a concentration of 0.100 mol / dm3
Calculate the mass of zinc iodide (ZnI_2) required to make 250 cm3 of this solution.

Relative atomic masses (*A*r): Zn = 65

_____

_____

_____

_____

_____

_____

Mass = _____ g

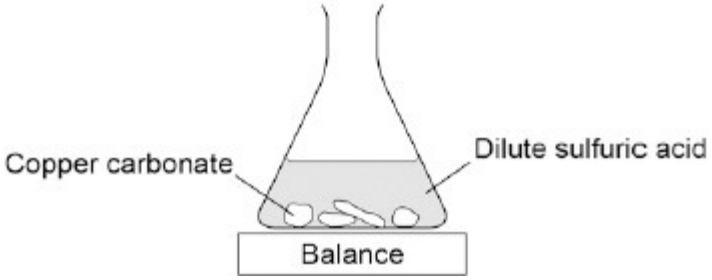
(3)

(Total 14 marks)

Q9.

A student investigated the reaction of copper carbonate with dilute sulfuric acid.

The student used the apparatus shown in the figure below.



(a) Complete the state symbols in the equation.

$CuCO_3$ (___) + H_2SO_4 (aq) → $CuSO_4$ (aq) + H_2O (___) + CO_2 (g)

(2)

(b) Why did the balance reading decrease during the reaction?

Tick one box.

The copper carbonate broke down. ☐

A salt was produced in the reaction. ☐

A gas was lost from the flask. ☐

Water was produced in the reaction. ☐

(1)

(c) Describe a safe method for making pure crystals of copper sulfate from copper carbonate and dilute sulfuric acid. Use the information in the figure above to help you.

In your method you should name all of the apparatus you will use.

_____

_____

_____

_____

_____

_____

_____

_____

_____

_____

(6)

(d) The percentage atom economy for a reaction is calculated using:

$$\frac{\text{Relative formula mass of desired product from equation}}{\text{Sum of relative formula masses of all reactants from equation}} \times 100$$

The equation for the reaction of copper carbonate and sulfuric acid is:

$CuCO_3 + H_2SO_4 \rightarrow CuSO_4 + H_2O + CO_2$

Relative formula masses : $CuCO_3$ = 123.5; H_2SO_4 = 98.0; $CuSO_4$ = 159.5 Calculate the percentage atom economy for making copper sulfate from
copper carbonate.

_____

_____

_____

_____

_____

Atom economy = _____ %

(3)

(e) Give one reason why is it important for the percentage atom economy of a reaction to be as high as possible.

_____

_____

(1)

(Total 13 marks)

Q10.

A student investigated the reactions of copper carbonate and copper oxide with dilute hydrochloric acid.

In both reactions one of the products is copper chloride.

(a) Describe how a sample of copper chloride crystals could be made from copper carbonate and dilute hydrochloric acid.

_____

_____

_____

_____

_____

_____

_____

_____

(4)

(b) A student wanted to make 11.0 g of copper chloride.

The equation for the reaction is:

$$CuCO3 + 2HCl \rightarrow CuCl2 + H2O + CO2$$

Relative atomic masses, *Ar*: H = 1; C = 12; O = 16; Cl = 35.5; Cu = 63.5

Calculate the mass of copper carbonate the student should react with dilute hydrochloric acid to make 11.0 g of copper chloride.

_____

_____

_____

_____

Mass of copper carbonate = _____ g

(4)

(c) The percentage yield of copper chloride was 79.1 %. Calculate the mass of copper chloride the student actually produced.

_____

_____

Actual mass of copper chloride produced = _____ g

(2)

(d) Look at the equations for the two reactions:

Reaction 1 $CuCO_3(s) + 2HCl(aq) \rightarrow CuCl_2(aq) + H_2O(l) + CO_2(g)$

Reaction 2 $CuO(s) + 2HCl(aq) \rightarrow CuCl_2(aq) + H_2O(l)$

Reactive formula masses: CuO = 79.5; HCl = 36.5; $CuCl_2$ = 134.5; H_2O = 18

The percentage atom economy for a reaction is calculated using:

$$\frac{\text{Relative formula mass of desired product from equation}}{\text{Sum of relative formula masses of all reactants from equation}} \times 100$$

Calculate the percentage atom economy for Reaction 2.

_____

_____

_____

_____

_____

_____

Percentage atom economy = _____ %

(3)

(e) The atom economy for Reaction 1 is 68.45 %.
Compare the atom economies of the two reactions for making copper chloride.

Give a reason for the difference.

_____

_____

(1)

(Total 14 marks)

Q11.

(a) Nitrogen and hydrogen are passed over iron to produce ammonia in the Haber Process.

Balance the equation for the reaction.

$$N_2 + H_2 \rightarrow NH_3$$

(1)

(b) What is iron used for in the Haber process?

Tick one box.

catalyst ☐

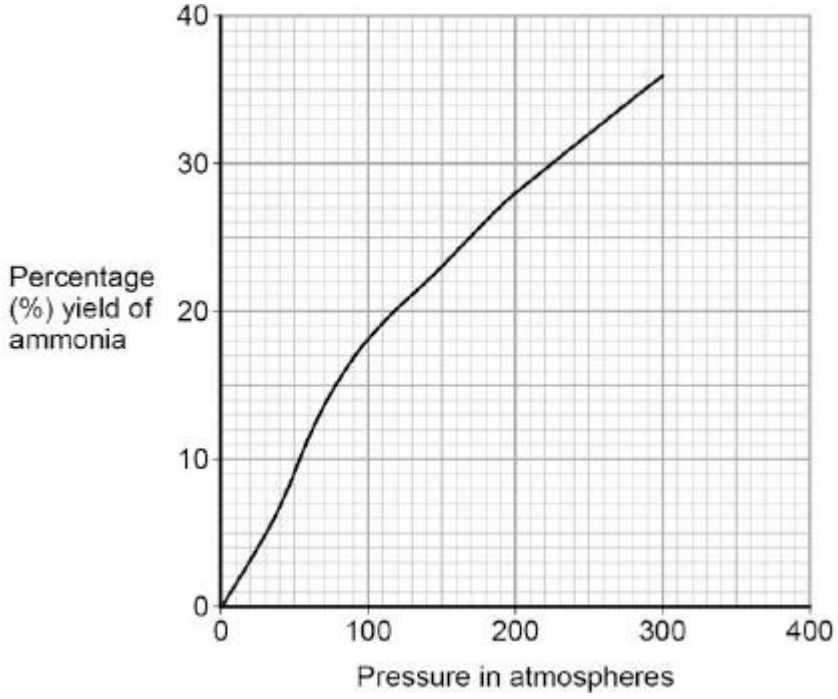
fuel ☐

monomer ☐

reactant ☐

(1)

(c) The figure below shows how the percentage yield of ammonia changes with pressure.



Describe the trend shown in the figure above.

_____

_____

(1)

(d) Use the figure above to determine the difference in percentage yield of ammonia at 150 atmospheres pressure and 250 atmospheres pressure.

Difference in percentage yield of ammonia = _____ %

(2)

(Total 5 marks)